

國立清華大學 命題紙

八十七學年度 化學系(所) 化學應用化學組碩士班研究生入學考試

科目 綜合化學 科號 0501 0601 共 9 頁第 1 頁 *請在試卷【答案卷】內作答

單選題，每題二分，不倒扣

1. A sample is brought into a laboratory and mixed with an equal volume of a preservative solution. For analysis a 5.00-mL sample is diluted to 100 mL, and the concentration of chloride ions in the diluted solution is found to be 3.0×10^{-3} M. What is the chloride concentration of the sample?
 (A) 6.0×10^{-3} M (B) 1.5×10^{-4} M (C) 7.5×10^{-5} M
 (D) 1.2×10^{-1} M

2. The relative uncertainty in the answer to the calculation $\frac{23.00 \times 1.68}{0.1416 \times 0.1332}$ is
 (A) 10 (B) 5×10^{-3} (C) 0.01 (D) 0.0001

3. A 35.25-mL sample is needed. The best piece of glassware to use is
 (A) a buret (B) a beaker (C) a graduated cylinder
 (D) a volumetric flask

4. A buffer with a pH of 10.0 is needed. Which of the following should be used?
 (A) acetic acid with a K_a of 1.8×10^{-5}
 (B) ammonia with a K_b of 1.8×10^{-5}
 (C) nitrous acid with a K_a of 7.1×10^{-4}
 (D) H_2PO_4^- and PO_4^{3-} with a K_a of 4.5×10^{-13}

5. An indicator has a K_a of 6.4×10^{-6} , the conjugate acid is red, and the conjugate base is yellow. At what pH will the solution be red?
 (A) 5.2 (B) 5.5 (C) 4.0 (D) 4.7

6. If 50.0 mL of a 0.0134 M HCl solution is mixed with 24.0 mL of a 0.0250 M NaOH solution, what is the pH of the final mixture?
 (A) 1.87 (B) 12.40 (C) 5.29 (D) 3.02

八十七學年度 化學系(所) 化學應用化學組碩士班研究生入學考試

科目 綜合化學 科號 0501 共 9 頁第 2 頁 *請在試卷【答案卷】內作答
科號 0601

7. A solution containing HF is titrated with KOH. At the end point of the titration the solution contains
(A) equal amounts of HF and KOH (B) H_2O , H^+ , OH^- , K^+ , F^- , and HF
(C) K^+ and F^- (D) KF and H_2O , and H_2O
8. What is the molar solubility of silver chromate? (K_{sp} for $\text{Ag}_2\text{CrO}_4 = 1.2 \times 10^{-12}$)
(A) 1.1×10^{-6} (B) 1.1×10^{-4} (C) 6.7×10^{-5} (D) 5.5×10^{-7}
9. In which of the following is electromagnetic radiation listed correctly in increasing order of wavelength (lowest wavelength first)?
(A) Ultraviolet < visible < infrared
(B) Visible < microwave < infrared
(C) X rays < infrared < ultraviolet
(D) Microwaves < visible < infrared
10. What is the major cause of the deviation of the Beer's law for the measurement of the acetic acid absorbance?
(A) the dimer chemical reaction (B) the wavelength band width chosen
(C) the slit width of the monochromator
(D) the background stray light
11. For the reduction of MnO_4^- to Mn^{2+} , the correct form of the Nernst equation is .
(A) $E = E^\circ + \frac{0.0591}{3} \log\left(\frac{[\text{Mn}^{2+}]}{[\text{MnO}_4^-][\text{H}^+]^8}\right)$
(B) $E = E^\circ + \frac{0.0591}{5} \log\left(\frac{[\text{Mn}^{2+}]}{[\text{MnO}_4^-][\text{H}^+]^8}\right)$
(C) $E^\circ = E + \frac{0.0591}{5} \log\left(\frac{a_{(\text{Mn}^{2+})}}{a_{(\text{MnO}_4^-)} \cdot a_{(\text{H}^+)}^8}\right)$
(D) $E = E^\circ + \frac{0.0591}{5} \log\left(\frac{a_{(\text{Mn}^{2+})}}{a_{(\text{MnO}_4^-)} \cdot a_{(\text{H}^+)}^8}\right)$

八十七學年度 化學系(所) 化學應用化學組碩士班研究生入學考試
 科目 綜合化學 科號 0501 共 9 頁第 3 頁 *請在試卷【答案卷】內作答

12. The K_{sp} of AgCl is 1.0×10^{-10} , and the K_{sp} of AgI is 8.3×10^{-17} . A solution is 0.100 M in I^- and Cl^- . What is the molarity of iodide ions when AgCl just starts to precipitate?
 (A) 1.0×10^{-5} (B) 9.1×10^{-9} (C) 8.3×10^{-7}
 (D) 8.3×10^{-8}
13. What kind of the information will be provided by the IR spectrum?
 (A) Electronic transition (B) Nuclear spin
 (C) Nuclear vibration (D) Electron spin
14. Which of the followings is a zero order reaction?
 (A) Thermal isomerization of cis-stilbene to trans-stilbene
 (B) enzyme oxidation of glucose to gluconic acid.
 (C) decay of triplet excited C_{60} to ground state.
 (D) decay of radioactivity of ^{60}Co .
 (E) none of the above.
15. For a first order chemical reaction $A \xrightarrow{k_1} B$, the concentration of reactant A can be best described by
 (A) $A = A_0 - kt$ (B) $A = A_0 e^{-k_1 t}$ (C) $1/A = 1/A_0 - kt$
 (D) $A = A_0 - e^{-k_1 t}$ (E) none of the above.
16. At constant T & P, the entropy change of the universe ΔS_{univ} for a spontaneous reaction,
 (A) $\Delta S_{univ} > 0$. (B) $\Delta S_{univ} = 0$ (C) $\Delta S_{univ} < 0$.
 (D) ΔS_{univ} can not be determined. (E) none of the above.
17. Which of the following equations is not correct?
 (A) $dG = VdP - SdT$ (B) $\mathcal{E} = \mathcal{E}^0 - (nF/RT) \ln(Q)$ (C) $(\partial S/\partial P)_T = -(\partial V/\partial T)_P$
 (D) $dP/dT = \Delta H_m/T\Delta V_m$ (E) none of the above.

八十七學年度	化 學	系 (所)	化 學 應用化學	組碩士班研究生入學考試
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科目 綜合化學 科號 0501 共 9 頁第 4 頁 *請在試卷【答題卷】內作答

18. Using the relationship $\ln(K) = -\Delta H^\circ/RT + \Delta S^\circ/R$. For an endothermic reaction at equilibrium, increasing the temperature will
 - (A) shift the equilibrium to right
 - (B) shift the equilibrium to left.
 - (C) reaction does not shift to either left or right.
 - (D) can not be determined.
 - (E) none of the above.
19. Which of the followings absorbs microwave?
 - (A) electronic transitions
 - (B) vibrational transitions
 - (C) rotational transitions.
 - (D) free electron transitions from β to α spin state.
 - (E) nuclear transitions from α to β spin state.
20. When a substance is photochemically excited to the electronic excited state, which of the properties of the molecules might be changed
 - (A) redox potential
 - (B) dipole moment
 - (C) the color of a solution containing the excited molecules.
 - (D) the pK_a value.
 - (E) all of the above.
21. 1 eV is equivalent to about
 - (A) 15 kcal/mol
 - (B) 23 kcal/mol
 - (C) 30 kcal/mol
 - (D) 42 kJ/mol
 - (E) none of the above
22. A mole of photon with 400 nm wavelength carries
 - (A) 40 kcal/mol
 - (B) 50 kcal/mol
 - (C) 60 kcal/mol
 - (D) 70 kcal/mol
 - (E) 80 kcal/mol.
23. In a closed container, there are equal moles of N_2 and O_2 . Which of the following description is not correct?
 - (A) N_2 and O_2 molecules have the same velocity.
 - (B) N_2 and O_2 molecules have the same kinetic energy.
 - (C) The partial pressure produced by N_2 and O_2 molecules are the same.

八十七學年度 化學系(所) 化學應用化學 組碩士班研究生入學考試

科目 綜合化學 科號 0501 共 9 頁第 5 頁 *請在試卷【答案卷】內作答

- (D) the numbers of molecules for N_2 and O_2 are the same.
 (E) all the above are correct.
24. Of the following, an important precaution to be observed in the storage of metallic sodium is to
 (A) leave the container uncovered.
 (B) store the sodium in kerosene.
 (C) store the sodium in water.
 (D) use an opaque container.
 (E) store the sodium in ethanol.
25. In the first transition series, the paramagnetism is due almost entirely to the unpaired spins, being approximately equal to $\mu = 2 \times \text{square root } [S(S+1)]$ magnetons where S is the total spin. On this basis, the Cu^+ ion has
 (A) zero magnetons (B) 3.88 magnetons (C) 2.83 magnetons.
 (D) 6.18 magnetons (E) 1.41 magnetons.
26. Which of the following compound is most likely to crystallize in the crystal system called *pervoskites*?
 (A) $MgCl_2$ (B) $BaTiO_3$ (C) $CaCl_2$ (D) Nb_2O_5 (E) $MgAl_2O_4$.
27. Which compound is central symmetrical, i.e. possessing a center of inversion?
 (A) XeF_4 (B) SF_4 (C) H_2O (D) CF_4 (E) PCl_5 .
28. Which of the Period 4 elements that forms neutral metal carbonyls with five carbonyl ligands.
 (A) silicon (B) germanium (C) nickel (D) iron (E) tungsten.
29. Of the following complexes, the one with the largest value of the crystal field splitting, Δ_o is
 (A) $[Fe(H_2O)_6]^{2+}$ (B) $[Ru(H_2O)_6]^{2+}$ (C) $[Ru(NH_3)_6]^{3+}$

八十七學年度 化學系(所) 化學應用化學組碩士班研究生入學考試

綜合化學

科號 0501
0601

共 9 頁第 6 頁 *請在試卷【答案卷】內作答

(D) $[\text{Os}(\text{CN})_6]^{3-}$.

30. The Jahn-Teller theorem states that for a molecule in an electrically degenerate state, distortion must occur to lower the symmetry, remove the degeneracy, and lower the energy. The transition metal complex that will not display the Jahn-Teller effect is
(A) $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ (B) $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ (C) $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ (D) $[\text{CoF}_6]^{3-}$.
31. Which of the following description is consistent with the properties of solid BN.
(A) A bluish-white, lustrous solid melting at 327°C ; the solid is soft and malleable.
(B) A white solid melting at 772°C ; the solid is an electrical nonconductor but dissolves in water to give a conducting solution.
(C) A yellowish-green solid melting at 172°C .
(D) A very hard, colorless substance melting at about 3000°C .
32. What is the bond-order between the nitrogen atom and the sulfur atom of the remarkable molecule, thiazyl trifluoride, NSF_3 ? The N-S distance determined by diffraction analyses is $1.416\text{\AA}$.
(A) single (B) double (C) one and half (D) triple (E) quadruple.
33. Which of the following amines shows the highest affinity with Lewis acid BF_3 in gas phase.
(A) NMe_3 (B) NHMe_2 (C) NH_2Me (D) NH_3 (E) NEt_3 .
34. What is the most convenient technique to provide the structural information of a volatile inorganic molecule, such as WF_6 in gaseous phase?
(A) X-ray diffraction method
(B) Electron diffraction method
(C) Extended X-ray absorption fine structure, EXAFS method
(D) Neutron diffraction method
(E) ^{19}F NMR spectroscopy.

八十七學年度 化學系(所) 化學系 組碩士班研究生入學考試
 科目 綜合化學 科號 0501 共 9 頁第 7 頁 *請在試卷【答案卷】內作答

35. In acidic solutions, soaps are converted to
 (A) insoluble salts (B) fatty acids (C) detergents (D) esters (E) glycerol.
36. Which of the following compound is most likely to crystallize in a structure in which the positive ions pack in the tetrahedral holes in a closest-packed array of negative atoms?
 (A) Li_2S (B) Li_3P (C) Cs_3P (D) CsCl (E) NaCl .
37. Which dinuclear compounds listed below possesses a metal-metal quadruple bond?
 (A) $\text{W}_2(\text{OMe})_6$ (B) $\text{Mo}_2(\text{MeCO}_2)_4$, $\text{MeCO}_2 = \text{acetate}$
 (C) $\text{Cp}_2\text{Re}_2(\text{CO})_4$, $\text{Cp} = \text{C}_5\text{H}_5$ (D) $\text{Cp}_2\text{W}_2(\text{CO})_6$.
38. Which of the following boron cluster complexes can be considered to adopt the so-called *nido*-configuration?
 (A) B_3H_8^- (B) B_4H_{10} (C) $\text{B}_{10}\text{C}_2\text{H}_{12}$ (D) $\text{B}_9\text{C}_2\text{H}_{11}^{2-}$.
39. Which of the following correctly ranks the cycloalkanes in order of increasing ring strain per methylene?
 A. cyclopropane < cyclobutane < cyclohexane < cyclopentane
 B. cyclohexane < cyclopentane < cyclobutane < cyclopropane
 C. cyclohexane < cyclobutane < cyclopentane < cyclopropane
 D. cyclopentane < cyclopropane < cyclobutane < cyclohexane
 E. cyclopropane < cyclopentane < cyclobutane < cyclohexane
40. Which of the following statements correctly pertains to a pair of enantiomers?
 A. They rotate the plane of polarized light by exactly the same amount and in opposite directions.
 B. They rotate the plane of polarized light by differing amounts and in opposite directions.
 C. They rotate the plane of polarized light by differing amounts and in the same direction.
 D. They have different melting points.
 E. They have the same melting point, but they have different boiling points.

國立清華大學 命題紙

八十七學年度	化 學	系 (所)	化 學 應用化學	組碩士班研究生入學考試
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科目 綜合化學 科號 0501 共 9 頁第 8 頁 *請在試卷【答案卷】內作答

41. Which of the following compounds will undergo an S_N2 reaction most readily?
- A. $(CH_3)_3CCH_2I$
 - B. $(CH_3)_3CCl$
 - C. $(CH_3)_2CHI$
 - D. $(CH_3)_2CHCH_2CH_2CH_2Cl$
 - E. $(CH_3)_2CHCH_2CH_2CH_2I$
42. Using Saytzeff's rule, choose the most stable alkene among the following.
- A. 1-methylcyclohexene
 - B. 3-methylcyclohexene
 - C. 4-methylcyclohexene
 - D. They are all of equal stability according to Saytzeff's rule.
43. Both (E)- and (Z)-3-hexene can be treated with O_3 in the presence of a platinum catalyst. How are the products from these two reactions related to each other?
- A. The (E)- and (Z)-isomers generate the same products in exactly the same amounts.
 - B. The (E)- and (Z)-isomers generate the same products but in differing amounts.
 - C. The products of the two isomers are related as diastereomers.
 - D. The products of the two isomers are related as enantiomers.
 - E. The products of the two isomers are related as structural isomers.
44. Which of the following alcohols undergoes dehydration upon heating with concentrated H_2SO_4 with no rearrangement of the carbon skeleton?
- A. 2-methyl-3-hexanol
 - B. 3-methyl-3-pentanol
 - C. 3,3-dimethyl-2-pentanol
 - D. 2-methyl-2-phenylpropanol
 - E. both a and b
45. A nucleus with an _____ atomic number or an _____ mass number has a nuclear spin that can be observed by the NMR spectrometer.
- A. odd, odd
 - B. odd, even
 - C. even, odd
 - D. even, even

國 立 清 華 大 學 命 題 紙

	八十七學年度	化 學	系 (所)	化 學 應用化學	組碩士班研究生入學考試
科目	綜合化學	科號	0501 0601	共 9 頁	第 9 頁 *請在試卷【答案卷】內作答

46. When 1,2-dibromobutane is heated at 200°C in the presence of molten KOH, what is the major organic product?

- A. 1-bromo-1-butyne
- B. 1-bromo-2-butyne
- C. 1-butyne
- D. 2-butyne
- E. 1-butene

47. Which of the following compounds is the most reactive dienophile in a Diels-Alder reaction with 1,3-butadiene?

- A. $\text{CH}_2=\text{CHOCH}_3$
- B. $\text{CH}_2=\text{CHCHO}$
- C. $\text{CH}_3\text{CH}=\text{CHCH}_3$
- D. $(\text{CH}_3)_2\text{C}=\text{CH}_2$
- E. $\text{CH}_2=\text{CH}_2$

48. In the UV-visible spectra of the following compounds, in which does λ_{max} appear at the highest wavelength?

- A. 1-phenylpropene
- B. 3-phenylpropene
- C. benzene
- D. n-propylbenzene
- E. isopropylbenzene

49. In electrophilic aromatic substitution reactions the nitro group is:

- A. a m-director since it destabilizes the meta sigma complex more than the ortho, para.
- B. a m-director since it destabilizes the meta sigma complex less than the ortho, para.
- C. an o,p-director since it stabilizes the ortho, para sigma complex more than the meta.
- D. an o,p-director since it stabilizes the ortho, para sigma complex less than the meta.
- E. none of the above.

50. When a carbonyl is part of a conjugated π -network, the C=O stretch:

- A. has a higher frequency than in a nonconjugated system.
- B. has a lower frequency than in a nonconjugated system.
- C. always occurs at 1710 cm^{-1} .
- D. occurs around 2700 cm^{-1} .
- E. cannot be distinguished from the C=O stretch in a nonconjugated system.

對數表 $y = \log_{10} x$

x											表尾差								
	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
10	0000	0043	0086	0128	0170	0212	0253	0294	0334	0374	4	8	12	17	21	25	29	33	37
11	0414	0453	0492	0531	0569	0607	0645	0682	0719	0755	4	8	11	15	19	23	26	30	34
12	0792	0828	0864	0899	0934	0969	1004	1038	1072	1106	3	7	10	14	17	21	24	28	31
13	1139	1173	1206	1239	1271	1303	1335	1367	1399	1430	3	6	10	13	16	19	23	26	29
14	1461	1492	1523	1553	1584	1614	1644	1673	1703	1732	3	6	9	12	15	18	21	24	27
15	1761	1790	1818	1847	1875	1903	1931	1959	1987	2014	3	6	8	11	14	17	20	22	25
16	2041	2068	2095	2122	2148	2175	2201	2227	2253	2279	3	5	8	11	13	16	18	21	24
17	2304	2330	2355	2380	2405	2430	2455	2480	2404	2529	2	5	7	10	12	15	17	20	22
18	2553	2577	2601	2625	2648	2672	2695	2718	2742	2765	2	5	7	9	12	14	16	19	21
19	2788	2810	2833	2856	2878	2900	2923	2945	2967	2989	2	4	7	9	11	13	16	18	20
20	3010	3032	3054	3075	3096	3118	3139	3160	3181	3201	2	4	6	8	11	13	15	17	19
21	3222	3243	3263	3284	3304	3324	3345	3365	3385	3404	2	4	6	8	10	12	14	16	18
22	3424	3444	3464	3483	3502	3522	3541	3560	3579	3598	2	4	6	8	10	12	14	15	17
23	3617	3636	3655	3674	3692	3711	3729	3747	3766	3784	2	4	6	7	9	11	13	15	17
24	3802	3820	3838	3856	3874	3892	3909	3927	3945	3962	2	4	5	7	9	11	12	14	16
25	3979	3997	4014	4031	4048	4065	4082	4099	4116	4133	2	3	5	7	9	10	12	14	15
26	4150	4166	4183	4200	4216	4232	4249	4265	4281	4298	2	3	5	7	8	10	11	13	15
27	4314	4330	4346	4362	4378	4393	4409	4425	4440	4456	2	3	5	6	8	9	11	13	14
28	4472	4487	4502	4518	4533	4548	4564	4579	4594	4609	2	3	5	6	8	9	11	12	14
29	4624	4639	4654	4669	4683	4698	4713	4728	4742	4757	1	3	4	6	7	9	10	12	13
30	4771	4786	4800	4814	4829	4843	4857	4871	4886	4900	1	3	4	6	7	9	10	11	13
31	4914	4928	4942	4955	4969	4983	4997	5011	5024	5038	1	3	4	6	7	8	10	11	12
32	5051	5065	5079	5092	5105	5119	5132	5145	5159	5172	1	3	4	5	7	8	9	11	12
33	5185	5198	5211	5224	5237	5250	5263	5276	5289	5302	1	3	4	5	6	8	9	10	12
34	5315	5328	5340	5353	5366	5378	5391	5403	5416	5428	1	3	4	5	6	8	9	10	11
35	5441	5453	5465	5478	5490	5502	5514	5527	5539	5551	1	2	4	5	6	7	9	10	11
36	5563	5575	5587	5599	5611	5623	5635	5647	5658	5670	1	2	4	5	6	7	8	10	11
37	5682	5694	5705	5717	5729	5740	5752	5763	5775	5786	1	2	3	5	6	7	8	9	10
38	5798	5809	5821	5832	5843	5855	5866	5877	5888	5899	1	2	3	5	6	7	8	9	10
39	5911	5922	5933	5944	5955	5966	5977	5988	5999	6010	1	2	3	4	5	7	8	9	10
40	6021	6031	6042	6053	6064	6075	6085	6096	6107	6117	1	2	3	4	5	6	8	9	10
41	6128	6138	6149	6160	6170	6180	6191	6201	6212	6222	1	2	3	4	5	6	7	8	9
42	6232	6243	6253	6263	6274	6284	6294	6304	6314	6325	1	2	3	4	5	6	7	8	9
43	6335	6345	6355	6365	6375	6385	6395	6405	6415	6425	1	2	3	4	5	6	7	8	9
44	6435	6444	6454	6464	6474	6484	6493	6503	6513	6522	1	2	3	4	5	6	7	8	9
45	6532	6542	6551	6561	6571	6580	6590	6599	6609	6618	1	2	3	4	5	6	7	8	9
46	6628	6637	6646	6656	6665	6675	6684	6693	6702	6712	1	2	3	4	5	6	7	7	8
47	6721	6730	6739	6749	6758	6767	6776	6785	6794	6803	1	2	3	4	5	5	6	7	8
48	6812	6821	6830	6839	6848	6857	6866	6875	6884	6893	1	2	3	4	5	5	6	7	8
49	6902	6911	6920	6928	6937	6946	6955	6964	6972	6981	1	2	3	4	4	5	6	7	8
50	6990	6998	7007	7016	7024	7033	7042	7050	7059	7067	1	2	3	3	4	5	6	7	8
51	7076	7084	7093	7101	7110	7118	7126	7135	7143	7152	1	2	3	3	4	5	6	7	8
52	7160	7168	7177	7185	7193	7202	7210	7218	7226	7235	1	2	2	3	4	5	6	7	7
53	7243	7251	7259	7267	7275	7284	7292	7300	7308	7316	1	2	2	3	4	5	6	6	7
54	7324	7332	7340	7348	7356	7364	7372	7380	7388	7396	1	2	2	3	4	5	6	6	7

x											表 尾 差								
	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
55	7404	7412	7419	7427	7435	7443	7451	7459	7466	7474	1	2	2	3	4	5	5	6	7
56	7482	7490	7497	7505	7513	7520	7528	7536	7543	7551	1	2	2	3	4	5	5	6	7
57	7559	7566	7574	7582	7589	7597	7604	7612	7619	7627	1	2	2	3	4	5	5	6	7
58	7634	7642	7649	7657	7664	7672	7679	7686	7694	7701	1	2	2	3	4	4	5	6	7
59	7709	7716	7723	7731	7738	7745	7752	7760	7767	7774	1	1	2	3	4	4	5	6	7
60	7782	7789	7796	7803	7810	7818	7825	7832	7839	7846	1	1	2	3	4	4	5	6	6
61	7853	7860	7868	7875	7882	7889	7896	7903	7910	7917	1	1	2	3	4	4	5	6	6
62	7924	7931	7938	7945	7952	7959	7966	7973	7980	7987	1	1	2	3	3	4	5	6	6
63	7993	8000	8007	8014	8021	8028	8035	8041	8048	8055	1	1	2	3	3	4	5	5	6
64	8062	8069	8075	8082	8089	8096	8102	8109	8116	8122	1	1	2	3	3	4	5	5	6
65	8129	8136	8142	8149	8156	8162	8169	8176	8182	8189	1	1	2	3	3	4	5	5	6
66	8195	8202	8209	8215	8222	8228	8235	8241	8248	8254	1	1	2	3	3	4	5	5	6
67	8261	8267	8274	8280	8287	8293	8299	8306	8312	8319	1	1	2	3	3	4	5	5	6
68	8325	8331	8338	8344	8351	8357	8363	8370	8376	8382	1	1	2	3	3	4	4	5	6
69	8388	8395	8401	8407	8414	8420	8426	8432	8439	8445	1	1	2	2	3	4	4	5	6
70	8451	8457	8463	8470	8476	8482	8488	8494	8500	8506	1	1	2	2	3	4	4	5	6
71	8513	8519	8525	8531	8537	8543	8549	8555	8561	8567	1	1	2	2	3	4	4	5	5
72	8573	8579	8585	8591	8597	8603	8609	8615	8621	8627	1	1	2	2	3	4	4	5	5
73	8633	8639	8645	8651	8657	8663	8669	8675	8681	8686	1	1	2	2	3	4	4	5	5
74	8692	8698	8704	8710	8716	8722	8727	8733	8739	8745	1	1	2	2	3	4	4	5	5
75	8751	8756	8762	8768	8774	8779	8785	8791	8797	8802	1	1	2	2	3	3	4	5	5
76	8808	8814	8820	8825	8831	8837	8842	8848	8854	8859	1	1	2	2	3	3	4	5	5
77	8865	8871	8876	8882	8887	8893	8899	8904	8910	8915	1	1	2	2	3	3	4	4	5
78	8921	8927	8932	8938	8943	8949	8954	8960	8965	8971	1	1	2	2	3	3	4	4	5
79	8976	8982	8987	8993	8998	9004	9009	9015	9020	9025	1	1	2	2	3	3	4	4	5
80	9031	9036	9042	9047	9053	9058	9063	9069	9074	9079	1	1	2	2	3	3	4	4	5
81	9085	9090	9096	9101	9106	9112	9117	9122	9128	9133	1	1	2	2	3	3	4	4	5
82	9138	9143	9149	9154	9159	9165	9170	9175	9180	9186	1	1	2	2	3	3	4	5	5
83	9191	9196	9201	9206	9212	9217	9222	9227	9232	9238	1	1	2	2	3	3	4	5	5
84	9243	9248	9253	9258	9263	9269	9274	9279	9284	9289	1	1	2	2	3	3	4	5	5
85	9294	9299	9304	9309	9315	9320	9325	9330	9335	9340	1	1	2	2	3	3	4	4	5
86	9345	9350	9355	9360	9365	9370	9375	9380	9385	9390	1	1	2	2	3	3	4	4	5
87	9395	9400	9405	9410	9415	9420	9425	9430	9435	9440	0	1	1	2	2	3	3	4	4
88	9445	9450	9455	9460	9465	9469	9474	9479	9484	9489	0	1	1	2	2	3	3	4	4
89	9494	9499	9504	9509	9513	9518	9523	9528	9533	9538	0	1	1	2	2	3	3	4	4
90	9542	9547	9552	9557	9562	9566	9571	9576	9581	9586	0	1	1	2	2	3	3	4	4
91	9590	9595	9600	9605	9609	9614	9619	9624	9628	9633	0	1	1	2	2	3	3	4	4
92	9638	9643	9647	9652	9657	9661	9666	9671	9675	9680	0	1	1	2	2	3	3	4	4
93	9685	9689	9694	9699	9703	9708	9713	9717	9722	9727	0	1	1	2	2	3	3	4	4
94	9731	9736	9741	9745	9750	9754	9759	9763	9768	9773	0	1	1	2	2	3	3	4	4
95	9777	9782	9786	9791	9795	9800	9805	9809	9814	9818	0	1	1	2	2	3	3	4	4
96	9823	9827	9832	9836	9841	9845	9850	9854	9859	9863	0	1	1	2	2	3	3	4	4
97	9868	9872	9877	9881	9886	9890	9894	9899	9903	9908	0	1	1	2	2	3	3	4	4
98	9912	9917	9921	9926	9930	9934	9939	9943	9948	9952	0	1	1	2	2	3	3	4	4
99	9956	9961	9965	9969	9974	9978	9983	9987	9991	9996	0	1	1	2	2	3	3	3	4
x	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9